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heat capacity, C

Heat brought to a system to increase its temperature divided by that temperature increase. At constant volume $C_V = \left(\frac{\partial U}{\partial T}\right)_V$ at constant pressure $C_p = \left(\frac{\partial H}{\partial T}\right)_p$, where U is the internal energy and H the enthalpy of the system.

Source:
Green Book, 2nd ed., p. 48
PAC, 1996, *68*, 957 *(Glossary of terms in quantities and units in Clinical Chemistry (IUPAC-IFCC Recommendations 1996))* on page 975

