

stoichiometric mean molal (practical activity coefficient)

in electrochemistry

The practical activity coefficient of electrolyte B is given by

$$\gamma_{\pm} = \frac{a_{\pm}}{\sqrt{\nu_+^{\nu_+} \nu_-^{\nu_-} \frac{m_B}{m^+}}}$$

where a_{\pm} is the mean activity of B in solution, m_B is the molality of B, $m^+ = 1 \text{ mol kg}^{-1}$, ν_+ is the number of cations and ν_- the number of anions in the chosen group B which is taken as the electrolyte.

$$\nu = \nu_+ + \nu_-$$

Source:

PAC, 1974, 37, 499 (*Electrochemical nomenclature*) on page 510